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54 **Porous siliceous-containing gas enriching material and process of manufacture and use.**

57 A hollow, porous, silica-rich fiber, a process for producing same, and a process for enriching at least one gas from a gaseous mixture utilizing at least one fiber to result in gas separations with good permeability and good selectivity. The fiber is non-crystalline and has pore sizes having a range of 1 to around 50 Angstroms in diameter and a mean pore size of around less than 20 Angstroms in diameter and a filament diameter in the range of 1 to around 200 microns and a wall thickness in the range of around 1 to 30 microns. The fibers with a fine pore structure and thin walls are produced by forming hollow glass fibers hindered or fully phase-separable boron and/or alkali metal glass composition, by attenuation from melt at speeds in the range of 500 ft/min to around 20,000 ft/min. The phase-separable glass fibers are heat treated to yield limited phase-separated hollow glass fibers and these fibers are leached to extract acid and/or water soluble components. For fully phase-separable, hollow glass fibers, a limited heat treatment is used before leaching, where the heat treatment is at a temperature of around 300 to less than 480°C for a time of around 5 minutes to 24 hours. The hindered, phase-separable, hollow glass fibers are heat treated at a temperature in the range of about 400 to about 600°C. At least one hollow, porous, silica-rich fiber is used in a device suitable to contain a gaseous mixture and the permeate is withdrawn from the silica-rich fiber at the opposite side from that at which the gaseous mixture contacts the hollow, porous fiber.

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## POROUS SILICEOUS-CONTAINING GAS ENRICHING MATERIAL AND PROCESSES OF MANUFACTURE AND USE

The present invention is directed to a porous, siliceous material for separating or enriching one gas or condensible vapor from one or more other gases and/or condensible vapors in a gaseous mixture and the processes of making and using the porous material.

Sundry types of membranes enable the separation of one gas or condensible vapor from one or more other gases and/or vapors in a gas mixture. These types of membranes include: polymer membranes, glass membranes, and composite or multicomponent membranes. The application of these membranes to separating gases addresses the need to enrich, separate, or isolate one gas from another in areas such as: chemical processing, petroleum processing, oxygen replenishing or carbon dioxide removal from air in sealed vehicles such as airplanes, space vehicles and submarines.

One example of the usefulness of separation or enriching membranes in the chemical process industry is the separation of helium from methane. Helium is not present in nature in pure form but only as a component of gaseous mixtures, both naturally occurring as in natural gas and artificially occurring as in by-product gas mixtures. The most widely used commercial process for separating helium from natural gas involves low temperature fractionation. Since helium is normally found in natural gas in concentrations of around 400 parts per million to 8 volume percent of natural gas, this separation process requires cooling of extremely large quantities of gas to temperatures low enough to liquify all of the gases except helium. The size of such a plant and the expensive equipment involved provides an impetus for utilizing membranes to separate helium from methane. Other examples in the area of chemical or petroleum processing include: removal of acid gases from fuel gas sources and synthesis gas in the area of gasification of fossil fuels, separation of carbon dioxide in the presence of acid gases (secondary oil recovery), oxygen separation from chlorine in chlorine production are azeotrope breaking. Also the membranes find application in gas analysis by their incorporation into analytical gas analyzing instruments.

Materials that are candidates for membranes must meet both physical and chemical requirements as well as yielding good membrane requirements. The physical property requirements include: good performance in the areas of tensile strength, tear strength, abrasion resistance, flexibility, dimensional stability for a wide range of thermal and chemical conditions, stress crack resistance, strength to withstand compaction in high-pressure utilization, good toughness and morphological stability for a reasonable period of time. Some chemical property requirements involve good performance in such areas as stability to withstand temperature fluctuations, chemical resistance, and maintenance of morphology and microstructure even under severe pressure changes. In addition, the materials should exhibit good membrane properties of selectivity at useful fluxes with adequate permeability. The right combination of permeability and selective separation and the thinnest possible form available embodies a membrane that would be useful in the gas enrichment or separation areas.

The glass membranes in the form of porous thimbles in multicomponent membranes proved useful in the gas separation of: hydrogen from hydrogen sulfide, and helium from methane. Researchers at the National Chemical Laboratory for Industry in Japan reported the separation of a helium and carbon dioxide gas mixture by means of a porous glass membrane in the form of a thimble utilizing cocurrent flow in a temperature range of 296 to 947° (564.8°F - 1,736.6°F). Researchers at the State Scientific-Research Institute of Glass; State Scientific-Research Institute of the Nitrogen Industry in the Soviet Union reported silica-rich highly porous glasses, approximately 96% SiO<sub>2</sub> in the form of microporous glass tubes were useful as semi-permeable membranes. In this work, the unchanging permeability coefficient in numerous experiments indicated the free-molecule character of the gas flow through the glass. Also it has been shown that the high-purity, nonporous, silica glass membranes can purify helium from natural gas and other sources since this smallest of gas molecules (2 Angstroms) can pass through the membrane via defects in the glass structure. German researchers (British Patent No. 1,271,811) utilizing a porous substrate with a nonporous glass glazing were successful in separating helium from methane. In the aforementioned work with porous high-silica containing glasses, a heat treatable borosilicate glass composition was used which was known as Vycor glass compositions available from Corning Glass Works. This composition has an amount of boron oxide in the range of about 20 to 35 weight percent and an amount of soda of 4 to 12 weight percent with the balance being silica. These glass compositions are phase separable upon heat treatment and are leached to remove a majority of the boron oxide and alkali metal oxides to leave around 96 weight percent silica-rich material.

In United States Patent 4,042,359, the Vycor glass compositions is used in forming porous capillaries. The capillaries having an outer diameter in the range of 40 to 140 microns since the internal diameter ranges from 30 to 80 microns and the wall thickness ranges from 5 to 30 microns. These capillaries are heat treated at a temperature ranging from 500° to 650°C for a total of 5 minutes to 50 hours. These heat treated capillaries are acid leached to produce porous capillaries having pore sizes of 11 Angstroms up to diameters of 1000 Angstroms. Smaller pore diameters can be achieved by shrinking the porous capillary. Heat treatment at about 700 to 800°C effects shrinkage to produce pore sizes with diameters down to less than 5 Å. The porous capillaries are useful in reverse osmosis water desalination processes.

In the area of membranes used for gas enrichment or separation, additional enhancements of this technology await the development of thinner, yet stronger membranes with uses under wider temperature conditions and with improved chemical durability.

#### SUMMARY OF THE INVENTION

The present invention in its three aspects include:

a gas separation and/or enrichment material that is at least one porous, silica-rich, hollow, inorganic fiber; a process of making same; and a process of separating or enriching at least one gas or condensible vapor from a mixture of gases and/or condensible vapors.

The gas separation and/or enrichment material has particular form and dimensions, composition and fine pore morphology. The form is a porous, hollow fiber having outer diameters in the range of around 1 to around 200 micrometers, um, (microns) and wall thicknesses in the range of around less than 1 micron to around 30 microns. Suitable lengths for the hollow fibers are those that are effective to permit a gaseous and/or vaporous mixture to contact one surface while permitting an enriched gaseous or vaporous stream to be collected at a second opposite surface of the hollow porous fiber. The silica-rich composition of the porous, hollow fiber results from leaching the acid soluble components of limited phase-separated, glass fibers. The fine pore morphology relates to pores having pore sizes in diameters in the range from around 1 to 50 Angstroms with a mean average pore size in the range of around 5 to 30 Angstroms. The pore sizes and the average pore size have good uniformity throughout the thin walls of the hollow fibers. This type of hollow glass fiber results from standard heat treatment of hindered, phase-separable hollow glass fibers or from limited heat treatment of phase-separable borosilicate and alkali metal silicate glass composition. The former approach involves heat treated glass compositions, where the glass compositions have a reduced tendency to phase separated; for example, binary and tertiary glass systems having ions with lower ionic potential reduced ionic size or increased ionic field strength. A reduced tendency for phase separation occurs when the glass composition has present at least one oxide of Group IVB of the Periodic Table of Elements or alkali metal oxides in decreasing order of effectiveness for retarding phase separation: Cs > K > Na > Li and aluminum oxide or a mixture thereof. The latter approach involves heat-treated, phase-separable borosilicate glass compositions that were heat treated at a temperature in the range of about 300°C to less than 480°C for an effective period of time.

The process of making the hollow, porous fiber membrane involves: forming, limited-phase-separating, and leaching the heat-treated fibers. Formation of the hollow fibers occurs at attenuation rates of around 500 to 30,000 feet/min. The fiber issues from a bushing having orifices designed to deliver gas pressures sufficient to give the fiber and inner to outer diameter ratio of around 0.2 to around 0.95 and outer diameters in the range of 1 to around 5 micrometers and a wall thickness of around less than 1 to around 50 micrometers. The glass composition of the hollow fibers is phase-separable ranging from hindered, phase-separable to fully phase-separable siliceous-containing compositions with at least 20 weight percent leachables. The leachables are materials soluble in appropriate acids, (those other than hydrofluoric or phosphoric acid) and/or water soluble materials. The hollow glass fibers are heat treated at temperatures ranging from 300°C to 480°C for the fully phase-separable glass fibers to from 400 to 600°C for the hindered, phase-separable glass fibers. The heat-treated hollow glass fibers are leached with appropriate acids and/or water to extract those leachable materials. This produces the hollow, porous, silica-rich inorganic fiber membrane.

The process of separating gases and/or condensible vapors involves using the porous hollow fibers so that the gaseous mixture contacts one side of at least one of the porous hollow fibers, and enriched or separated gas (permeate) is removed from the opposite side. If the gaseous mixture contacts the exterior surface of the porous hollow fiber, the opposite side is the lumen or interior surface of the hollow fiber. The gaseous mixture is removed from the same side of the porous hollow fiber that is initially contacted. Either

the lumen or the exterior surface of the hollow glass fiber can be the side for initial contact by the gaseous mixture. The at least one porous hollow fiber is housed in a suitable apparatus to contain the gaseous mixture and provide for removal of the gaseous mixture at one location and of the permeate from a separate location connected to the opposite side of the porous hollow fiber. The conditions of separating or enriching include thermodynamic driving forces and temperatures ranging from sub-ambient to elevated temperatures below the softening point of the silica rich fibers.

## DETAILED DESCRIPTION OF THE INVENTION

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In the following description and in the claims, the following terms have the described meanings.

The term "gas" includes any condensible vapors, where the vapors are condensible at conditions other than those used in the feed gas for enrichment or separation.

15 The term "perm selective" means permeable to different extents to different molecular species under equal driving force.

The term "ultrafiltered type" means transport is partially by a porous mechanism but membrane penetrant interactions are also important.

20 The term "permeability factor" represents the steady state rate of gas transmission through a membrane. Values for permeability factors are not ordinarily normalized for membrane thickness. For homogenous membranes, the permeability factor is inversely proportional to the sample thickness. When the thickness of the active part of the membrane is not known, e.g., in asymmetric membranes, the permeability factor is still a valid permeability characterization. The permeability factor values determine the value-in-use of the membrane in permeation devices. The derivation of the permeability factor equation is as follows: the volume of gas transmitted through a membrane is directly proportional to the area, time and pressure of the permeation test as:

permeability factor = volume divided by area x time x pressure

The units selected for volume, area, time and pressure are cm<sup>3</sup> (STP), cm<sup>2</sup>, second and cm of mercury, respectively.

30 The term "gas separation" means separations ranging from 100 percent separation to enrichment of a gas in any concentration over and above that of the original gaseous mixture.

The term "separation factor" or "selectivity" for a membrane and for a given gaseous mixture including gases A, B, ....N is defined as the ratio of the permeability constant ( $P_a$ ) of the membrane for gas (A) to the permeability constant ( $P_{others}$ ) of the membrane for the other gases in the mixture. The separation factor is also equal to the ratio of the permeability factor ( $P_a/1$ ) of a membrane of thickness "1" for gas "a" of a gas mixture to the permeability factors of the same membrane to the other gases ( $P_{others}/1$ ) wherein the permeability constant or permeability for a given gas is the volume of gas at standard temperature and pressure (STP) which passes through a membrane per square centimeter of surface area, per second, for a partial pressure drop of 1 centimeter of mercury across the membrane per unit of thickness and is expressed as  $P = \text{cm}^3 \cdot \text{cm} / \text{cm}^2 \cdot \text{second} \cdot \text{cm of mercury}$ . In practice, the separation factor with respect to a given pair of gases in a binary gaseous mixture for a given membrane can be determined by employing numerous techniques which provide sufficient information for calculation of permeability constants or permeability factor for each of the pair of gases in the binary mixture. Several of the many techniques available for determining permeability constants, permeabilities and separation factors are disclosed by the work entitled "Techniques of Chemistry", Vol. VII, Membranes & Separations, by Hwang et al., John Wiley & Sons, 1975, herein incorporated by reference at Chapter 12, pages 296-322.

45 The terminology "porous separation membrane" relates to membranes which may have continuous porous channels for gaseous flow that communicate between the interior surface and the exterior surface, or which have numerous fine pores in the range of 1 to 50 Angstroms, where a majority of the pores interconnect. Also the term refers to porous materials which have numerous pores in the range of 1 to 50 Angstroms and which may have interconnecting pores for gaseous flow between the interior surface and the exterior surface. In other words, the membrane has considerable internal void volume related to continuous porosity from one side.

55 The term "selectivity" is defined as the ratio of the rate of passage of the more readily passed component of a gaseous mixture to the rate of passage of the less readily passed components. Selectivity may be obtained directly by contacting a membrane with a known mixture of gases and analyzing the permeate. Alternatively, a first approximation of the selectivity can be obtained by setting up the ratio of the rates of passage of the two or more components determined separately on the same membrane. The rates may be expressed as GcB units.

Permeability can be measured by the variable pressure method or the variable volume method, both methods are well known to those skilled in the art.

The term "phase separated" refers to glass fiber compositions that form immiscible phases upon heat treatment. After the appropriate heat treatment, the size of the composition fluctuations are greater than the size of a critical nucleus of 20 Angstroms. Mechanisms for describing the formation of new phases include the classical nucleation and growth mechanisms. Here, there is a sharp boundary between phases with a measurable interfacial energy and the new stable phase is defined by a critical nucleus size. For glasses, critical nucleus size has been estimated to be larger than 20Å as shown in "Direct Measurements of Homogeneous Nucleation Rates in a Glass-Forming System", by J. J. Hammel, Jl. Chem. Physics, Vol 46, No. 6, pp. 2234-2244 (1967), hereby incorporated by reference. Glasses can also phase separate by a mechanism of "Spinodal Decomposition" in which small fluctuations in composition grow rapidly when the glasses are in the spinodal region of a miscibility gap. Here, initially, there is no sharp boundary or interfacial energy between phases, and the inhomogeneities (or composition fluctuations) are points of maximum concentration of one component that gradually move to maximum concentration of the second component, e.g., SiO<sub>2</sub> to B<sub>2</sub>O<sub>3</sub> in a borosilicate glass. This type structure is also found above the miscibility gap where there are similar fluctuations in composition even in a single phase glass melt. Therefore, glass structure in the initial stages of spinodal decomposition cannot be distinguished from the structure of a single phase glass. Sharp boundaries with measurable interfacial energies are a necessary requirement for two-phase systems (e.g., phase separated glass). These conditions are not approached in spinodal separating systems until the phases are at least the size of a critical nucleus of around 20 Angstroms.

It is believed without limiting the scope of the invention that the porous, hollow, silica-rich fiber membranes of the present invention acts as a membrane, where the gaseous separation is due to more than size exclusion and penetrant interactions but also includes interactions with the porous walls to give good permeability and good selectivity.

For a better understanding of the invention, the term "extractable" for a group of components refers to metal oxides and associated materials which are leachable from borosilicate, phase-separable glass fibers by acids other than hydrofluoric and phosphoric acids. These latter acids cannot be used since they attack silica. Also the term "non-extractable" for the group of components refers to silica and metal oxides of Group IVB of the Periodic Chart. These materials are not leachable from the glass by acids other than hydrofluoric and phosphoric acids. Also the terms "associated material" refers to an interconnected phase of interaction products of the components of the extractable group or the non-extractable group or of the components from both groups because of their proximity to each other in the glass fibers. Nonexclusive examples of associated materials include alkali metal borates, alkali metal aluminates, other interaction products with aluminum oxide and the like.

The porous, hollow, silica-rich fibers of the present invention are prepared from a formulated fiberizable batch composition to result in, pore generating, hindered or fully, phase-separable, glass fibers. The batch composition is formulated to enable the resulting glass fibers to have a balance of the two groups of components, i.e., the extractable group including associated materials and the non-extractable group. Typical batch materials known to those skilled in the art can be used and calculated by known methods to produce the glass compositions with the components of the two groups.

In the pore generating glass fiber composition, the components and affiliated materials of the extractable group are present in an amount of at least 20 to around 60 volume percent and include: one or more boron-containing materials; alkali metal oxides (R<sub>2</sub>O); alkaline earth metal oxides like bivalent oxides (CaO, MgO); trivalent oxides like Al<sub>2</sub>O<sub>3</sub>, and Fe<sub>2</sub>O<sub>3</sub> and oxides such as TiO<sub>2</sub>, SnO<sub>2</sub>, and P<sub>2</sub>O<sub>5</sub>. In addition, trace amounts of materials usually present in trace amounts in glass fibers can also be present like fluorine. The total amount of the components of the extractable group in the glass fibers is preferably at least 30 volume percent of the total glass composition. The presence of the R<sub>2</sub>O and/or B<sub>2</sub>O<sub>3</sub> materials also serve as a fluxes in producing the glass fibers.

In the extractable group, the concentration of boron-containing material, boron oxides and/or anhydride and affiliated materials having boron, should not be too great that the softening point of the glass fibers decreases to cause the fibers to stick together during any heat treatment. This tackiness complicates the maintenance of individual fiber identify and, for discrete lengthy fibers, the maintenance of any near parallel alignment during subsequent processing steps.

The components of the non-extractable group include the siliceous material such as silica and any refractory glass modifiers like the tetravalent oxides of zirconium and/or titanium and/or hafnium.

Any phase-separable glass compositions (those that separate into phases upon heat treatment) within the range of the aforelisted compositions having little, if any, refractory glass can be used in the present invention. This includes those having high concentrations of boron-containing materials, i.e., around 40 to around 60 weight percent of the total glass composition. The hindered phase separable glass compositions (those that phase separate upon heat treatment but that contain at least one refractory glass modifiers) also are useable in the present invention. Both the phase-separable and the hindered phase-separable glass fibers are used in phase-separated form resulting from heat treatment. Examples of the hindered, phase-separable glass compositions include the borosilicate or alkali metal silicate that have refractory glass modifiers such as metal oxides from Group IVB and Periodic Table, aluminum oxide and alkali metal oxides in decreasing order of hinderance of phase separation  $Cs > K > Na > Li$  oxides and mixtures thereof.

A suitable hindered phase-separable alkali metal silica glass composition for forming hollow fibers has around 29 weight percent sodium oxide ( $Na_2O$ ), 62 weight percent silica ( $SiO_2$ ), and 9 weight percent zirconium oxide ( $ZrO_2$ ). For these types of glass compositions, the amount of alkali metal oxide  $R_2O$  can range from about 20 to about 40 weight percent and the amount of metal oxide of the Group IV metal can range from about 1 to about 20, preferably less than 12 weight percent  $ZrO_2$ , with the remainder being silica.

When the content of boron-containing material is 0 to slightly greater than zero, the amount of alkali metal oxides ( $R_2O$ ) can be a major component in the extractable group. When the amount of boron-containing compound is in the range of less than 20 weight percent of the total resulting glass composition, one or more of the other aforelisted inorganic oxide components are present in the extractable group. When the amount of boron-containing material is around 35 weight percent and up to 60 weight percent of the total resulting glass composition, the boron-containing material components constitute the majority and the other inorganic oxide components constitute the minority of the extractable group.

Regarding the other components in the glass fiber that can be considered in the extractable group, the aluminum oxide can be present in an amount in the range of 0 to about 15 weight percent of glass composition. This amount is generally less for higher amounts of boron-containing materials in the glass fibers and larger for lower amounts of boron-containing materials in the glass fibers. The amount of  $R_2O$  components range from less than one weight percent up to around 15 weight percent, when higher amounts of boron-containing components are present. Also there can be present, especially with lower amounts of boron-containing components, calcium oxide ( $CaO$ ) and magnesium oxide ( $MgO$ ). The total amount of these components can be in the range of 0 to about 30 weight percent of the glass composition.

The amount of siliceous material should not be less than around 30 and preferably not less than 40 weight percent of the total glass composition. Generally, the siliceous material is less than around 80, preferably less than around 70 weight percent of the total glass composition. The metal oxides such as zirconium and/or titanium and/or hafnium can be present in amounts from 1 to about 20 weight percent of the glass composition. Preferably, zirconium oxide is present in an amount of up to about 8 weight percent. Since these oxides have good acid insolubility, they are present with silica in the porous hollow fiber. These metal oxides not only render porous fibers with good alkaline stability, but also enable substitution of the tetravalent oxides of zirconium and/or titanium and/or hafnium for one or more of the components of the extractable group. Their presence not only results in controlling porosity (by decreasing porosity without altering the amount of silica), but it also results in more alkaline stable, porous, silica-rich fibers.

The fully phase-separable borosilicate glass compositions include those pioneered by Hood and Nordberg for Corning Glass Company as described in U.S. Patents 2,106,744; 2,215,039; 2,221,709; 2,286,275; 2,461,841 and 2,494,259 and the article entitled "Properties of Some Vycor-brand Glasses", M. E. Nordberg, Journal of the American Ceramic Society, Vol. 27, No. 10, pgs. 299-305, all of which are incorporated herein by reference. Generally, the Vycor-brand glass composition has:  $B_2O_3$  in an amount of 20-35 weight percent, alkali metal oxide in an amount of 4-12 weight percent and silica in an amount of not less than 56 and not more than 75 weight percent. Also aluminum oxide may be present in an amount of not more than around 12 weight percent.

Another particularly useful, fully phase-separable glass composition for increased porosity is a glass having a low silica and high-borate content. Generally these glasses have amounts in weight percent of silica at 30 to 50 percent by weight, boric oxide at 40 to 55 percent by weight, alkali metal oxide at 5 to 15 percent by weight, aluminum oxide from 0 to 4 weight percent and zirconium oxide about 1 to about 4 weight percent.

The glass batch compositions are melted in a furnace at temperatures and times to obtain a fiberizable viscosity for the molten glass without devitrification. Generally, the batch is heated to 2000°F (1093°C) to 3000°F (1649°C) for 1 to about 6 hours or longer. The molten glass is attenuated from the orifices of a bushing located on a forehearth connected to the furnace. The bushing has tubes aligned and associated

with the orifices and connected to a supply of gas at a superatmospheric pressure to allow for a continuous flow of gas to the vicinity of the orifice. The flow of gas can be uniform to produce continuous glass fibers, or can be intermittent to produce intermittent hollow glass fibers. A further description of the production of hollow and intermittent glass fibers is given for a direct melt system in U.S. Patents 3,268,313; 3,421,873 and 3,526,487, all hereby incorporated by reference. In the alternative, a pre-mixed feed of raw materials in batch form or as glass marbles can be used in making glass fibers in an indirect melt separation.

Preferably, the production of hollow glass fibers of the instant invention results in good concentricity of the central lumen of the fibers for two basic reasons.

First, the bushing tip is firmly affixed to the faceplate of the bushing. The aligned tube through which the gases are introduced to provide the central lumen of the glass fibers as they are being formed is also rigidly affixed to the bushing faceplate through a bracket member which forms a truss that prevents the tube from moving in any direction with respect to the faceplate itself. Thus, any warping of the faceplate during operation and consequent movement of the bushing tips carries with it an associated similar movement of the tube or conduit so that the gas introduced through the gas tube is always being introduced to the molten glass emanating from channels in the bushing tip at the same location. The established lumen of the resulting fibers is thus readily maintained at its formed diameter as is the outside diameter of the fiber. This provides for uniform K values (outer diameter/inner diameter) obtained in the strands herein produced. Preferably, the air tube extends beyond the tip of the orifice of the bushing to assist in producing hollow glass fibers with uniform K values.

Secondly, it has been found that by introducing the air stream into the glass at the exit point of the tip and at a larger diameter than heretofore used, any characteristic bulging of the cone of the molten glass is avoided. Further, a more stable attenuation process is achieved, and uniform concentric holes are provided in the glass fibers formed as the molten glass emanates from the bushing tip. This occurs because the diameters of the air stream and glass stream at the tip are in approximately the same proportion as the desired end product. Once again, the lumen is centrally located and is constantly in the same position (regardless of whether or not the bushing faceplate distorts) due to the firm connection between the bracket member, the bushing faceplate and the tube. Glass can flow freely into an area above the bushing tips in all instances, where the tubes are being held, since the bracket member is completely open in between the spaces between tabs and also through the holes located in the top of the bracket member. Welds between the tube and the bracket member and the solid connection or button formed by the side arm tabs of the brackets in the hole provided in the faceplate for the tabs provide a rigid, secure, truss-type attachment so that there is no movement of the tubes when faceplate distorts over time.

The attenuation of the glass fibers is conducted by mechanical means through winding or chopping, where in winding the fibers are grouped into a strand and wound onto a forming tube situated on a rotating mandrel of a winding apparatus. Any other method of forming and attenuating hollow fibers as known by those skilled in the art can also be used. As the fibers are attenuated at speeds on the order of around 500 to around 50,000 feet/minute into ambient temperature air they are cooled, and they can be treated with water or cooled air for additional cooling. The hollow glass fibers are collected separately or are gathered into one or more strands, usually by means of a gathering shoe. The fibers or strands are wound onto a rotating drum-type winder having a forming tube to produce a forming package. The collet on which the forming package rides usually rotates at high speeds to collect the strand or strands into the forming package. Such speeds generally can be around 4,400 or 6,000 revolutions per minute which continues until the winder is slowed to a stop and a forming package is removed. An example of the gathering and collecting of the glass fibers into a forming package is disclosed in U.S. Patent 4,071,339 (Griffiths) and U.S. Patent 4,049,411 (Long and Dent) where attenuation speeds of from about 2,000 to 20,000 feet per minute are achieved, both patents are hereby incorporated by reference.

It is believed without limiting the present invention, that the fast attenuation speeds assist in providing an open glass structure in the fiber to facilitate extraction of acid and/or water extractable components. Also, it is believed that rapid cooling of the fibers assists in producing an open network structure which allows leachable components to be extracted in reasonable time periods.

The strands of the hollow glass fibers can comprise any number of fibers known to those skilled in the art. The proper combination of bushing tip size and attenuation speed results in hollow fibers with diameters larger than around 200 microns or more, but preferably from around 1 to around 50 microns. Fibers with diameters larger than around 200 microns may be difficult to wind and may be formed in a manner similar to capillaries or tubes as is known in the art. The fibers can have a K factor of up to around 0.96 but preferably in the range of around 0.5 up to about 0.96. Best results are obtained when the hollowness of the glass fiber comprises around 10 to around 70 percent of the volume of the glass fibers. Fibers having more than 70 percent of their volume that is shallow can be unstable because of thin walls especially when the

fibers are leached. Finer fibers with outer diameters of less than 40 microns also are subject to such volume percentages of hollowness. Preferably, the hollowness is found in such a manner that the wall thickness is in the range of around 1 to around 30 microns and preferably from 1 to around 10 and most preferably up to about 5 microns.

5 The hollow glass fibers and/or strands that are collected into the forms of multilayered package, either forming packages or roving packages, or into the forms of chopped fibers or strands, chopped or continuous fibrous or strand mats or batts are treated for pore generation through limited phase-separation of the glass and subsequent leaching of extractables with appropriate acid and/or water. The continuous fibers or strands may be removed from the collection packages by cutting transverse to or parallel with the  
10 axis of the package or by rewinding onto larger diameter drums or can remain in the package, mat, or batt form for the generation of pores. Preferably the strands are cut from one or more multilayered packages by making one or more transverse cuts through the layers of the package. The length of the cut hollow glass fibers can be varied by varying the diameter of the forming package during winding of the hollow glass fibers or by rewinding the hollow glass fibers from the forming package onto a smaller or larger diameter  
15 package. The many layers of the hollow glass fibers which are removed from the package can be laid flat on a supporting surface. The supporting surface can be a plate or tray or moving conveyor belt. Generally, the discrete lengths of hollow glass fibers obtained by this approach can range from about 1 inch to around 25 inches. Any other method known to those skilled in the art for removing the glass fibers from the multilayered package can be employed. For example, the fibers can be unwound from the package and  
20 disposed as chopped strand or continuous strand onto another supporting surface or holder or rotating drum. Preferably, the discrete lengths of glass fibers can range from about 0.25 inch (64 cm) to around 70 inches (180 cm) and most preferably only up to around 25 inches (64 cm).

Before the pores are generated in the hollow glass fibers as fibers or strands through extraction, any sizing composition present on the fibers can be removed through a solvent wash such as a water wash to  
25 remove a water soluble sizing composition. It is preferred not to remove any sizing composition from the glass fibers, since the sizing composition may protect the heat-treated, phase-separated glass fibers during leaching to produce fine pores.

To achieve less than full phase-separation i.e., limited phase separation, for the hollow glass fibers the conditions of the heat treatment of the fibers depends on the composition of the fibers. Heat treatment of  
30 fibers with hindered-phase-separable compositions initiates the formation of an interconnected phase of non-extractable components and the formation of an interconnected phase extractable component. The heat treatment results in a mean pore diameter in the desired range for the porous hollow fibers. These hollow glass fibers are heat treated in a furnace usually at a temperature in the range of about 400 to about 500°C. The temperature assists in controlling the porosity of the glass fibers as taught for glass in U.S. Patent  
35 3,758,284 hereby incorporated by reference. Generally the heat treatment is for a period of time from about 10 minutes to several days depending somewhat on the concentration for glass modifiers in the composition. Longer times require lower temperatures, while shorter times require higher temperatures within the given temperature range. Most preferably, for porous, hollow, silica-rich fibers having a pore volume in the range of 0.5 to 1.2 cc/gm and pore diameters in the range of around 10 to 50 Angstroms with an average  
40 pore diameter of less than 20 Å, the fibers are heat treated at 470°C for 6 hours to phase separate the hindered phase-separable glass into the water and/or acid leachable components and the acid insoluble components of the siliceous group and Group IVB metal oxides. Afterwards, the fibers are cooled to ambient temperatures. For hollow glass fibers with fully phase-separable glass compositions, the heat treatment is limited. Heat treatment is performed in the aforescribed manner, except the temperature is  
45 limited to a range of about 300°C to less than 480°C.

The extraction for the heat treated fibers generates pores by acid and/or water leaching of the glass fibers. The specific type of leaching depends on the phase-separated, glass composition. The generated pores have a diameter across some portion of the opening of the pore, whether the pore is circular,  
50 elliptical, cylindrical or asymmetrical in shape, to yield the desired mean pore diameter. The mean pore diameter of the pores generated can have a broad or narrow distribution.

The acid and/or water leaching can be conducted at a temperature ranging from sub-ambient to an elevated temperature of the boiling point of the acid or water. Preferably the temperature is in the range of ambient to around 95°C. The acid is any organic acid or any inorganic acid other than hydrofluoric and phosphoric acids. The concentrations of the acids can be in the range of around 0.1 Normal to 12 Normal  
55 for an acid having a pKa similar to hydrochloric acid. Acids with a higher pKa can be used in more dilute solutions, and acids with a lower pKa can be used in higher concentrations. Also the water and/or acid leaching solution may have ions like those of an alkali metal borate solution to reduce occurrences of stress cracking. The time of leaching depends on the specific composition, the temperature of leaching and the



concentration of the acid. Generally, the times range from around 5 minutes for water or acid leaching of fibers with high concentrations of boron-containing extractables. The time is also as long as 24 to 72 hours for heat-treated fibers with low concentrations of boron-containing extractables. Also it is preferred to have higher acid concentrations to reduce any attack on the nonextractable components of the glass fibers to maintain adequate strength of the porous hollow fibers. Maintenance of the concentration of the acid at a low pH or of the water at near neutral pH in the leaching bath assists in providing favorable leaching kinetics.

The porous, hollow, silica-rich fibers of the invention, have around at least 75 weight percent silica possibly with a minor amount (up to around 8 weight percent) of the Group IVB metal oxide. Although, in any case, trace amounts of other metal oxides of the remainder of the silica-rich fiber composition may be present. The fibers have a mean pore diameter which can be controlled to be in the range of about 10 to about 200 angstroms from the various aforementioned compositions and extraction operations. The lengths vary from long chopped fibers of 1 inch to 25 inches to near continuous fibers. The porous fibers have a pore volume in the range of about 0.15 to about 0.75 cc/gm and fiber diameters from preferably about 3 microns to about 50 microns. The pore diameter is in the range of around 1 to 200 angstroms, preferably 1 to 100 angstroms and most preferably 1 to 50 angstroms. The average pore diameter is less than 50 angstroms and preferably less than around 20 angstroms. The fibrous membrane has good uniformity of pore size, but can be further treated to result in asymmetric pore structures or morphology by heating to condense some pores or by selective leaching with acid or alkaline materials.

The hollow, porous, silica-rich fibers of the present invention can be used for gas separations in any gas separating apparatus known to those skilled in the art. For example, the gas separation apparatus utilized by those skilled in the art for separating gases with the use of hollow polymeric fibers can also be used with the porous, hollow, silica-rich fibers of the present invention. Any of these apparatus that provide contact between a gaseous mixture and either the exterior surface or the lumen of at least one of the hollow, porous, silica-rich fibers and that provide for separate removal of the gaseous mixture and the permeate (separated gas) can be used. Removal of the permeate is from the opposite side of the fiber from which the gas mixture contact the hollow, porous fiber. An example of an apparatus that can be used is that shown in the article entitled "Helium and Diffusion Separation", by K. B. McAfee, Jr., Bell Laboratories Record, Vol. 39, pg. 358, October, 1961, hereby incorporated by reference. An example of gases that can be separated from gaseous mixtures include: helium from natural gas or a 50/50 helium methane mixture; oxygen from an oxygen nitrogen mixture, carbon dioxide from a carbon dioxide methane mixture; oxygen from an oxygen chlorine mixture and acid gases from gaseous mixtures containing same.

### 35 PREFERRED EMBODIMENT OF THE INVENTION

The glass forming fiberizable pore generating batch composition is formulated by back calculations to result in hollow glass fiber compositions that undergo limited heat treatment for phase-separation into the components of the extractable and nonextractable group. The hollow glass fibers are formed by melting the batch at about 2600°F (1427°C) for around 3 hours and by conditioning the melted batch in the hollow fiber bushing melter at 2600°F (1427°C) for around one hour and by mechanically attenuating the fibers. The hollow fibers formed from hollow fiber bushings with extended air tubes have good concentricity to have a diameter in the range of about 3 microns to about 100 microns and most preferably about 10 microns to about 40 microns and to have a wall thickness of around 1 to about 30 microns, most preferably 1 to 5 microns. The hollow glass fibers are preferably sized with an aqueous chemical treating composition having at least a fiber lubricant and possibly a film forming polymer. The fibers are gathered into one or more strands and wound into a cylindrical forming package.

A plurality of undried forming packages are cut transverse to the package longitudinal axis so that all the layers of glass fibers can be removed from the package. These fibers are pulled taut and are laid straight on trays or in crucibles in a nearly parallel alignment, where the fibers usually have a discrete length of about 25 inches (63.5 cm). The fibers are heat treated at a temperature in the range of 300 to less than 480°C, where the temperature is increased from ambient to 400°C in around one hour with the fibers present. The heat treating lasts for around 6 hours.

The hollow glass fibers are acid leached with agitation, in an acid solution, such as 0.1 to about 12 Normal, preferably, about 2 to 3 Normal hydrochloric acid, at temperatures around 50°C to 60°C, for about 5 minutes to about 8 hours, preferably about 1 to about 2 hours. In the acid leach, the hollow glass fibers are immersed in the acid bath for a sufficient period of time, at a sufficient temperature to remove a substantial amount, if not all, of the acid soluble boron-containing compounds in the glass fibers. The fibers

typically may be submerged in the acid, when the acid is cool, and the temperature of the acid is increased to the elevated temperature. Nonexclusive examples of other suitable solutions of acids include nitric acid, and other acids that give soluble leachate products. The volume ratio of acid to glass fibers in the acid leaching step can be about 1 to about 100 volumes of acid of one volume of hollow glass fibers which will vary somewhat with the normality of the acid. The hollow, silica-rich fibers are removed from the acid leaching solution, water washed to a pH of around 5.5 to neutral in the wash water after which the fibers are dried, preferably in air drying at a temperature of ambient to around 50°C for around 10 minutes to about 24 hours. In addition, the fibers can be further dried by techniques known to those skilled in the art to result in uniform drying.

A plurality of the porous, hollow, silica-rich fibers may be placed in a pressurizable container having an inlet and two outlets. One end of each fiber at the same side of the container is sealed. The unsealed side of each fiber is located in one of the outlets of the container. This allows for flow of the gaseous mixture into the container so that the permeate passes through and exits the porous hollow fibers and exits the container via the appropriate outlet of the container. The unseparated gaseous mixture exits the container at the other outlet and can be recirculated to the inlet.

The invention along with alternative embodiments are illustrated further in the following examples.

#### Example 1

Approximately 50 grams of hollow glass fibers with an outer diameter (OD) equal to 40 microns and an internal diameter (ID) equal to 30 microns were formed using a single tip bushing having an extended air tube design. The fiber were wound directly onto a cardboard forming tube with no binder at 1750 rpm from a melt of 2095°F (1146°C). The air tube had an internal air pressure of 5.4 inches (13.7 cm) of water. The glass composition of the glass fibers was as follows:

	<u>Oxide</u>	<u>Weight Percent</u>
30	Silica	68.6
	Boron oxide ( $B_2O_3$ )	24.8
	Soda ( $Na_2O$ )	5.0
35	Aluminum oxide ( $Al_2O_3$ )	1.5
	Ferric oxide ( $Fe_2O_3$ )	0.3
40	Percent leachables	31.6

Approximately 6 grams of the fiber were cut away from the forming package as 2½ inch (6.35 cm) wide ring. The ring was pulled taut at the opposite sides of the circumference and folded twice length-wise and stuffed into a 100 milliliter porcelain crucible. The crucible was placed into a muffle furnace which was programmed to heat from room temperature to 450°C in one hour then hold at 450°C for 6 hours at which time the crucible was immediately removed to ambient conditions. A 6 inch long piece of the heat treated material, approximately 1 gram by weight, was leached for 30 minutes in 1000 milliliters of 3 Normal HCl contained in a Pyrex dish at a temperature of 55°C. The sample was then rinsed repeatedly with deionized water.

In a similar manner to that of Example 1, hollow glass fibers can be made from compositions of Examples 2 through 9 of Table 1.

Table 1

LEACHABLE FIBER GLASS COMPOSITION FOR MEMBRANES

Examples Oxide	<u>2</u>	<u>3</u>	<u>4</u>	<u>5</u>	<u>6</u>	<u>7</u>	<u>8</u>	<u>9</u>
SiO <sub>2</sub>	59.0	65.7	63.7	39.8	54.5	52.5	39.8	52.4
B <sub>2</sub> O <sub>3</sub>	27.6	21.7	19.0	47.1	34.2	36.2	50.7	36.1
Na <sub>2</sub> O	8.5	8.3	10.3	9.0	7.0	7.0	9.4	7.0
ZrO <sub>2</sub>	3.3	2.0	2.0	3.0	3.0	3.0	-----	3.0
Al <sub>2</sub> O <sub>3</sub>	1.2	2.2	2.2	1.1	1.1	1.1	-----	1.1
CaO	-----	2.7	-----	-----	-----	-----	-----	-----
Fe <sub>2</sub> O <sub>3</sub>	0.3	-----	-----	-----	-----	-----	-----	0.3
Trace Materials <sup>1</sup>	0.1	0.1	0.1	-----	0.2	0.2	0.1	0.1
% Leachables	36.1	32.2	34.0	57.2	42.3	44.3	60.1	44.6

<sup>1</sup>Trace materials include such oxides as: MgO, TiO<sub>2</sub>, K<sub>2</sub>O, Cr<sub>2</sub>O<sub>3</sub>, SrO and BaO and fluorine

In preparing porous hollow fiber membranes, the hollow glass fibers of Table 1 are subjected to a heat treatment similar to that of Example 1. An exception is that the fully phase-separable glass compositions of Example 8 can be heat treated at a lower temperature in the range of about 300°C to less than 480°C.

- 5 As an illustration, the hollow glass fibers having a composition of Example 2 of Table 1 are formed and made porous at the conditions listed in Table 1.

## Table 2 FORMATION OF HOLLOW POROUS FIBERS

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### Forming Conditions:

- Attenuation Speed (ft/min) 1750 rpm  
 Bushing temperature (°F)/(°C) 2095  
 15 Sizing None  
 Air Tip Pressure (inches of H<sub>2</sub>O) 5.4  
 Heat Treatment (°C/hr.) 450/6

### 20 Leaching Conditions:

- Acid/Normality HCl/3  
 Temperature/time (°C/hr) 55/0.5  
 Glass Weight/Acid Volume (2gm/ml) 1/1000

- 25 The Example 2 hollow, porous fiber was prepared according to the conditions listed in Table 2 in a manner similar to that for Example 1.

- The hollow glass fibers were prepared from a single tip bushing designed with an extended air tube. The fibers were wound directly onto a cardboard forming tube. The air tip pressure from the extended air tube was 5.4 inches of water. The fibers were then heat treated and leached in a manner similar to that for  
 30 Example 1.

- Several of the porous, hollow, silica-rich fibers of Example 2 were potted into a single fiber capillary test cell. The cells are generally made with a 1/8 inch stainless steel tube approximately 2 inches (10 mm) long. An epoxy seal covers the fiber end internal to the cell and fills a short length, roughly 0.5 mm, of the fiber lumen to prevent gas from entering at that point. A second epoxy seal fills the space between the wall of  
 35 the capillary tube and the hollow fiber for a length of approximately 10 mm to prevent escape of gas there, except for a small vent fiber or tube. When the capillary inlet is connected by fittings to a source of pressurized feed gas mixture, the components must permeate the exposed walls from the exterior to the interior of the hollow porous fiber, which is approximately 1 inch long, and travel the length of the lumen to escape as permeate flow from the exit end of the hollow porous silica-rich fiber. The capillary cell is  
 40 oriented vertically with the fiber exit at the upper end so that a collection cell filled with water can be placed over the extended tip of the porous hollow fiber.

- The collection cell was a glass capillary tube having a 6 mm outside diameter and a 4 mm inner diameter which was sealed at one end and open at the other. The glass capillary was filled to overflowing with water so that the surface tension of the liquid allowed the glass capillary tube to be turned upside down  
 45 over the extended hollow fiber from the capillary cell without any of the water dripping out. The permeate gas from the single fiber capillary test cell allowed for bubbling of permeate gas out of the exit end of the porous hollow fiber and the gas collected at the top end of the collection cell as a large bubble. The volumetric rate of water displacement was equated to permeate flux and a gas tight syringe was used to transfer a sample of the gas for analysis by gas chromatography. The permeability was calculated  
 50 according to the formula  $P/L = \text{flux (cm}^2/\text{seconds/fiber area)} \times \text{cm}^2 \times \text{pressure drop (cm mercury)}$ . The selectivity is the ratio of the permeability of the fastest permeating gas to that of the slower permeating gas.

Table 3 presents the data of the selectivities and permeabilities obtained for various gas separations for the porous hollow fibers of Example 2. Also Table 3 presents published data for gas separations by a commercial polysulfone fiber.

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Table 3

## SELECTIVITIES AND PERMEABILITY FOR POROUS HOLLOW FIBERS

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MEMBRANE MATERIAL	GAS MIXTURE	PRESSURE PSIG	SELECTIVITY (P <sub>fast</sub> /P <sub>slow</sub> )		ESTIMATED PERMEABILITY FACTOR (P/l)	
			Actual	Knudsen predict	$\frac{\text{cm}^3 \times 10}{\text{cm}^2 \cdot \text{sec} \cdot \text{Hg}}$	
Example from Table 2	He /CH <sub>4</sub> 50/50	300	999	2.0	He	
"	O <sub>2</sub> /N <sub>2</sub> 23/77		1.75	0.94	O <sub>2</sub>	0.07
(POLY- SULFONE)	H <sub>2</sub> /CH <sub>4</sub>	69-80		2.8	H <sub>2</sub>	57-75
AS DESCRIBED IN U.S. PAT 4,472,175 AT TABLE 12	O <sub>2</sub> /N <sub>2</sub> (21/78)		4.5	1.1	O <sub>2</sub>	6-9
"	CO <sub>2</sub> /CH <sub>4</sub> 50/50	30		1.6	CO <sub>2</sub>	24

In Table 3, the data were accumulated with a number of aforementioned single fiber capillary cells for each gas separation, and the average values are presented in Table 3. For example, the membrane material example from Table 2 in separating helium from the 50/50 mixture of helium and methane was tested in 6 single fiber capillary cells. The permeability factor P/L was calculated as follows:

P/L = permeate/flux [cm<sup>3</sup>/seconds(stp)]/membrane area (cm<sup>2</sup>) . delta p(cm mercury)

delta P<sub>He</sub> = (300 psi) 76 cm mercury/14.7 psi x 0.5 (for 50% helium in the gaseous mixture = 776 cm mercury)

membrane area for 3.8 cm fiber of 40 micron OD 30 micron ID gives an average radius of 17.5 microns and surface area is equal to 2 pi average radius L is equal to 0.042 cm<sup>2</sup>

the permeate flux is equal to (.99A) (54 microliters) stp/minutes x cm<sup>3</sup>/1000 microliters x minutes/60 seconds = 8.98 x 10<sup>-4</sup> cm<sup>3</sup> (stp/seconds)

permeability factor is equal to 8.98 x 10<sup>-4</sup> cm<sup>3</sup> (stp/seconds)/4.2 x 10<sup>-2</sup> cm<sup>3</sup> (776 cm mercury) = 2.8 x 10<sup>-5</sup> cm<sup>3</sup> (stp)/cm<sup>2</sup> x cm mercury is equal to 28 x 10<sup>-6</sup>

selectivity P<sub>He</sub>/P<sub>CH<sub>4</sub></sub> where the denominator (area x delta P) for each term in the ratio cancels since delta P is for helium and methane were equivalent in the 50/50 gaseous mixture resulting in selectivity equaling 99.8/0.2 = 499 for the example from Table 2.

For all of the separations of Table 3, the Knudsen separation factor was calculated based on the molecular weights of the gases being separated.

At the bottom of Table 3, there are presented published results for gas separations of a commercial polysulfone fiber.

From Table 3, the results show very good selectivity in the separation of the gases with a very good permeability factor, in some cases, better than the polysulfone by orders of magnitude and in other cases, comparable to the polysulfone membrane.

In accordance with the aforescribed disclosure of hollow, porous, high silica fibers with fine pores and thin walls for separating gases from gaseous mixtures, we claim the following as our invention.

## 5 Claims

1. A gas separation material, comprising:  
 at least one hollow, porous, non-crystalline, silica-rich fiber having pore sizes within the range of about 1 to around 50 Angstroms in diameter, a mean average pore size of around less than 30 Angstroms in diameter,  
 10 filament diameters in the range of 1 to around 200 microns and a wall thickness in the range of around 1 to around 30 microns, wherein the silica-rich fiber results from leaching of the acid or water extractable materials from at least one incompletely, phase-separated, hollow glass fiber selected from the group consisting of:  
 hindered, phase-separable, hollow glass fibers heat treated at a temperatures in the range of 400 to 600°C.  
 15 for about 0.5 to 12 hours and fully phase-separable boron and/or alkali metal silicate, hollow glass fibers heat treated at a temperature in the range of about 300° to less than 480°C. for a time in the range of about 0.5 to 12 hours where the borosilicate glass fibers have an amount of boron oxide from 20 to 60 weight percent, where the hollow glass fibers are formed by attenuation from melt at speeds in the range of at least 500 ft/min. to around 50,000 ft/min. to have filament diameters in the range of around 1 to around 200  
 20 microns and a wall thickness in the range of around 1 to around 30 microns.
2. Article of Claim 1, wherein the pore volume of the hollow, porous, silica-rich fiber is in the range of 400 up to around 600 meters<sup>2</sup> per gram.
3. Article of Claim 1, wherein the fiber diameter is less than 40 microns.
4. Article of Claim 1, wherein the fiber diameter is in the range of 10 to 30 microns.
- 25 5. Article of Claim 1, wherein the attenuation speed is in the range of 1,000 ft/min to around 5,000 ft/min.
6. Article of Claim 1, wherein the wall thickness is in the range of 1 to about 5 microns.
7. Article of Claim 1, wherein the mean pore size is less than 30 Angstroms in diameter.
8. Article of Claim 1, wherein the mean pore size is less than 20 Angstroms in diameter.
- 30 9. Article of Claim 1, wherein the glass fibers are heat treated at a temperature in the range of 400 to less than 480°C for a time in the range of about 1 minute to 10 hours before the extractable materials are removed.
10. Article of Claim 1, wherein the pore volume of the hollow, porous, silica-rich fiber is in the range of 400 up to around 600 meters<sup>2</sup> per gram.
- 35 11. Article of Claim 1, wherein the fiber diameter is less than 40 microns.
12. Article of Claim 1, wherein the fiber diameter is in the range of 10 to 30 microns.
13. Article of Claim 1, wherein the wall thickness is in the range of 1 to about 5.
14. Article of Claim 1, wherein the mean pore size is less than 20 Angstroms in diameter.
15. Gas separation material of Claim 1, wherein the hindered, phase-separable glass fibers are  
 40 borosilicate glass fibers having one or more oxides of a metal from Group IBV of the Periodic Table, and the fully phase-separable glass fibers are borosilicate glass fibers with an amount of boron oxide greater than 35 up to around 60 weight percent of the glass fibers, and alkali metal silicate glass fibers.
16. A process of enriching at least one gas from a mixture of gases utilizing a porous, silica-rich membrane material, comprising:  
 45 using at least one hollow porous silica-rich fiber that is non-crystalline and has pore sizes within the range of 1 to around 50 Angstroms in diameter, a mean or average pore size of around 5 to 50 Angstroms in diameter, filament diameter in the range of around 1 to 200 microns and wall thickness in the range of around 1 to 30 microns, wherein the silica-rich fiber results from extracting extractable components selected from the group consisting of acid extractable components and water extractable components or a mixture  
 50 thereof, from hollow glass fibers formed by attenuation from melt at speeds in the range of at least 500 ft/min to around 20,000 ft/min where the glass fibers have a composition selected from the group consisting of hindered, non-phase separable, hollow glass fibers heat treated at a temperature in the range of 400 to 600°C. for about 0.5 to 12 hours and phase-separable borosilicate, hollow, fully phase-separable glass fibers heat treated at a temperature in the range of about 400° to less than 480°C. for a time in the range of  
 55 about 0.5 to 12 hours where the borosilicate glass fibers have an amount of boron oxide from 20 to 35 weight percent.

17. Process of Claim 16, wherein using the porous silica-rich membrane material includes:  
 contacting the gaseous mixture with one surface of the hollow, porous, silica-rich fiber at temperature in the  
 range of subatmospheric to elevated temperatures and at a pressure in the range of about 1 to 5,000 psi;  
 and

5 removing from the vicinity of the opposite surface of the hollow porous silica-rich fiber from the contact  
 surface, a product enriched in at least one gas of said gaseous mixture.

18. Process of Claim 16, wherein helium is separated from the gaseous mixture selected from the  
 group consisting of helium and methane, natural gas.

19. Process of Claim 16, wherein oxygen is separated from a gaseous mixture selected from the group  
 10 consisting of oxygen and nitrogen, air, carbon dioxide containing air and oxygen and chlorine gaseous  
 mixtures.

20. Process of Claim 16, wherein acid gases are removed from acid gaseous containing gaseous  
 mixtures.

21. Process of Claim 16, wherein the gas separation fiber for separating at least one gas from a  
 15 gaseous mixture, has:

pore sizes within the range of 1 to around 50 Angstroms in diameter and a mean average pore size in the  
 range of around less than 20 Angstroms in diameter and a filament diameter in the range of around 1  
 micron to 100 microns and a wall thickness in the range of around 1 micron to 10 microns, and wherein the  
 silica-rich fibers result from extracting acid soluble components from the hollow glass fiber having a  
 20 hindered, phase-separable borosilicate glass fiber composition with one or more oxides of a metal of Group  
 IVB of the Periodic Table.

22. A process for producing a gas enriching membrane, comprising:

a) forming at least one hollow glass fiber having a composition selected from the group consisting of  
 extractable, hindered, phase-separable, boron and/or alkali metal silicates and extractable, fully phase-  
 25 separable boron and/or alkali metal silicates both having an amount of at least 20 weight percent  
 extractables by acids other than hydrofluoric or phosphoric acid, extractables or water extractables of at  
 least 20 weight percent, where the hollow glass fibers have an outer diameter in the range of 1 to 200  
 microns (um) and have a wall thickness in the range of around 1 to around 30 microns (um) and have an  
 inner to outer diameter ratio in the range of 0.5 to 0.95 and where the fibers are attenuated (formed) at a  
 30 speed in the range of 500 to 50,000 feet/min. with rapid cooling of the fibers,

b) heat treating the hollow glass fibers to result in limited phase-separated, hollow glass fibers,

c) leaching the heat treated, hollow glass fibers to remove extractables selected from the group  
 consisting of acid soluble with acids other than hydrofluoric and phosphoric acids and water soluble and  
 mixtures thereof to produce a porous, silica-rich, hollow fiber having pores with diameters from 1 to 50 A  
 35 and with a mean pore diameter of less than 30 A and with a pore volume in the range of around 0.15 to  
 around 0.75 cm<sup>3</sup>/gm.

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